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chemical equilibrium . Haber s solution lay in finding a suitable catalyst which would allow the mixture to come to equilibrium at a lower temperature where K

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chemical solution equilibrium  $A + 2B \rightleftharpoons X + Y$  equilibrium constant

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Chemical equilibrium has been reached in a reaction when the rate of the forward reaction is equal  $K_c$   
symbolizes the equilibrium constant in an aqueous solution,

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The concept of chemical equilibrium was The diagram above illustrates the point that a precipitate that is not one of the main species in the solution equilibrium

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 $[ZnCl_2] = 1 \cdot 10^{-10} M$   $K_{sp}$  of  $MgCO_3 = 6.8 \cdot 10^{-6}$   $ZnCO_3$

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